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5-Benzyl-5*H*-pyrido[3,2-*b*]indole

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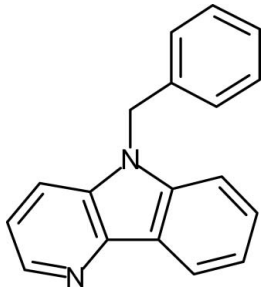
Key indicators: single-crystal X-ray study; T = 173 K; mean σ(C–C) = 0.004 Å; R factor = 0.055; wR factor = 0.149; data-to-parameter ratio = 17.3.

The title compound, $C_{18}H_{14}N_2$, was prepared by twofold Pd-catalyzed arylamination of a cyclic pyrido–benzo–iodolium salt. In the crystal, two molecules of 9-benzyl-δ-carboline form centrosymmetrical dimers with distances of 3.651 (2) Å between the centroids of the pyridine rings and 3.961 (2) Å between the centroids of the pyrrole and pyridine rings. The phenyl rings point to the other molecule in the dimer and the carboline core is essentially planar [maximum deviation of 0.027 (2) Å].

Related literature

For δ-Carboline, see: Subbaraju *et al.* (2004); Paulo *et al.* (2000); Chernyshev *et al.* (2001); Namjoshi *et al.* (2011); Qu *et al.* (2009); Masterova *et al.* (2008). For synthetic strategies to carbolines, see: Späth & Eiter (1940); Sakamoto *et al.* (1999); Franck *et al.* (2008). For the transition-metal-catalyzed synthesis of carbazoles, see: Letessier (2011); Nemkovich *et al.* (2009). For the transition-metal-catalyzed synthesis of carbolines, see: Nissen *et al.* (2011), Dassonneville *et al.* (2010). For β-carboline, see: Torreiles *et al.* (1985); Love (2006); Dassonneville *et al.* (2011); Nissen & Detert (2011). For the synthesis of the title compound, see: Letessier & Detert (2011).



Experimental

Crystal data

$C_{18}H_{14}N_2$
M_r = 258.1
Monoclinic, $P2_1/n$
a = 11.295 (4) Å
b = 10.482 (4) Å
c = 11.961 (4) Å
β = 110.387 (11)°
V = 1327.4 (8) Å^3
Z = 4
Mo $K\alpha$ radiation
μ = 0.08 mm^{-1}
T = 173 K
0.51 × 0.25 × 0.02 mm

Data collection

Bruker SMART CCD diffractometer
15877 measured reflections
3149 independent reflections
1444 reflections with $I > 2\sigma(I)$
R_{int} = 0.128

Refinement

$R[F^2 > 2\sigma(F^2)]$ = 0.055
$wR(F^2)$ = 0.149
S = 0.98
3149 reflections
182 parameters
H-atom parameters constrained
$\Delta\rho_{max}$ = 0.22 e Å^{-3}
$\Delta\rho_{min}$ = −0.21 e Å^{-3}

Data collection: *SMART* (Bruker, 2006); cell refinement: *SAINT* (Bruker, 2006); data reduction: *SAINT*; program(s) used to solve structure: *SIR97* (Altomare *et al.*, 1999); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *PLATON* (Spek, 2009); software used to prepare material for publication: *PLATON*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BT5607).

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supplementary materials

Acta Cryst. (2011). E**67**, o2341 [**doi:10.1107/S1600536811032107**]

5-Benzyl-5*H*-pyrido[3,2-*b*]indole

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Comment

The title compound was prepared as a part of a project focused on the transition metal catalyzed synthesis of carbazoles, see: Letessier (2011) and Nemkovich *et al.* (2009), carbolines, see: Nissen, Schollmeyer & Detert (2011) and related indolo-annulated heterocycles, see: Dassonneville *et al.* (2010). Whereas the β-carboline is the core of a large group of alkaloids (see: Torreiles *et al.* (1985); Love (2006)), only a few natural δ-carbolines are known. With Rh- or Ru-catalyzed [2 + 2+2] cycloadditions of alkynyl-ynamides we recently reported a new access to β- and γ-carbolines (Dassonneville *et al.*, 2011), Nissen & Detert (2011), but this approach is not suitable for the synthesis of the δ-isomers. These can now be prepared in a twofold Pd-catalyzed arylation of primary amines with cyclic pyrido-benzo iodolium salts. This unique 9-substituted δ-carboline crystallizes in form of centrosymmetrical dimers. The phenyl group, pointing in the direction of the second molecule, is nearly orthogonal to the essentially planar carboline core (maximal deviations of 0.027 (2) Å from the least square plane). Short distances of the centroid of a pyridine ring of one molecule to the centroid of the pyridine of its counterpart of 3.65 Å and to the pyrrole centroid of 3.96 Å indicate a π-π interaction between the heterocycles.

Experimental

A solution of 400 mg (0.93 mmol) of benzo[4,5]iodolo[3,2-*b*]pyridin-5-ium trifluoromethanesulfonate (Letessier & Detert, 2011) in dry toluene (10 ml) was deaerated in a Schlenk flask. Under argon, $Pd_2(dba)_3$ (34 mg, 0.04 mmol), Xantphos (64 mg, 0.11 mmol), and Cs_2CO_3 (850 mg, 2.61 mmol) were added. The mixture was stirred for 5 min at 300 K before benzyl amine (120 mg, 1.12 mmol) was added. After stirring for 15 h at 383 K, the mixture was cooled to ambient temparature, filtered through celite and concentrated. Purification by column chromatography (petroleum ether / ethyl acetate = 4 / 1) gave 156 mg (65%) of the title compound as colorless crystals with m. p. > 415 K.

Refinement

Hydrogen atoms attached to carbons were placed at calculated positions with C—H = 0.95 Å (aromatic) or 0.98–0.99 Å (sp^3 C-atom). All H atoms were refined in the riding-model approximation with isotropic displacement parameters (set at 1.2–1.5 times of the U_{eq} of the parent atom).

Figures

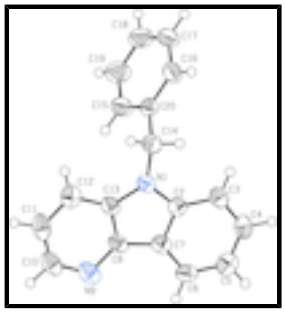


Fig. 1. View of compound **I**. Displacement ellipsoids are drawn at the 50% probability level.

5-Benzyl-5*H*-pyrido[3,2-*b*]indole

Crystal data

$C_{18}H_{14}N_2$	$F(000) = 544$
$M_r = 258.1$	$D_x = 1.293$ Mg m^{-3}
Monoclinic, $P2_1/n$	Melting point: 415 K
Hall symbol: -P 2yn	Mo $K\alpha$ radiation, $\lambda = 0.71073$ Å
$a = 11.295$ (4) Å	Cell parameters from 1161 reflections
$b = 10.482$ (4) Å	$\theta = 2.6$–$22.3°$
$c = 11.961$ (4) Å	$\mu = 0.08$ mm^{-1}
$\beta = 110.387$ (11)°	$T = 173$ K
$V = 1327.4$ (8) Å^3	Plate, colourless
$Z = 4$	0.51 × 0.25 × 0.02 mm

Data collection

Bruker SMART CCD diffractometer	1444 reflections with $I > 2\sigma(I)$
Radiation source: sealed Tube	$R_{int} = 0.128$
graphite	$\theta_{max} = 27.9°$, $\theta_{min} = 2.1°$
CCD scan	$h = -14 \rightarrow 14$
15877 measured reflections	$k = -13 \rightarrow 13$
3149 independent reflections	$l = -15 \rightarrow 15$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.055$	H-atom parameters constrained
$wR(F^2) = 0.149$	$w = 1/[\sigma^2(F_o^2) + (0.0558P)^2 + 0.2129P]$ where $P = (F_o^2 + 2F_c^2)/3$
$S = 0.98$	$(\Delta/\sigma)_{max} < 0.001$
3149 reflections	$\Delta\rho_{max} = 0.22$ e Å^{-3}

182 parameters	$\Delta\rho_{min} = -0.21$ e Å^{-3}
0 restraints	Extinction correction: *SHELXL97* (Sheldrick, 2008), $Fc^*=kFc[1+0.001xFc^2\lambda^3/\sin(2\theta)]^{-1/4}$
Primary atom site location: structure-invariant direct methods	Extinction coefficient: 0.016 (2)

Special details

Experimental. ^{1}H-NMR (400 MHz, $CDCl_3$): δ = 8.59 (dd, J = 4.7 Hz, J = 1.2 Hz, 1H, 2-H), 8.43 (d, J = 7.8 Hz, 1H, 9-H), 7.64 (dd, J = 8.2 Hz, J = 1.2 Hz, 1H, 6-H), 7.53 (td, J = 8.3 Hz, J = 1.2 Hz, 1H, 7-H), 7.42 (d, J = 8.2 Hz, 1H, 6-H), 7.26-7.34 (m, 5H, CH), 7.12 (m, 2H, CH), 5.52 (s, 2H, CH_2). ^{13}C-NMR (75 MHz, $CDCl_3$): δ = 141.9 (d, C-2), 141.8 (s, C-9b), 141.4 (s, C-5a), 136.5 (s, C-1), 134.0 (s, C-4a), 128.9 (d, C-2), 127.9 (d, C-7), 127.7 (d, C-4), 126.3 (d, C-3), 122.2 (s, C-9a), 120.9 (d, C-3), 120.1 (d, C-9), 120.0 (d, C-8), 115.8 (d, C-4), 109.2 (d, C-6), 46.5 (t, CH_2). IR (neat, ATR): ν = 1621 (w), 1588 (w), 1482 (m), 1451 (m), 1412 (s), 1334 (m), 1318 (s), 1242 (w), 1193 (m), 1115 (w), 1012 (w), 913 (w), 845 (m), 781 (s), 742 (vs), 730 (vs), 721 (vs), 695 (s)cm^{-1}. FD-MS: m/z = 258.1 $[C_{18}H_{14}N_2]^+$.

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted *R*-factor *wR* and goodness of fit *S* are based on F^2, conventional *R*-factors *R* are based on *F*, with *F* set to zero for negative F^2. The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating *R*-factors(gt) *etc*. and is not relevant to the choice of reflections for refinement. *R*-factors based on F^2 are statistically about twice as large as those based on *F*, and *R*- factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (Å^2)

	x	*y*	*z*	U_{iso}*/U_{eq}
N1	0.74211 (17)	0.47469 (18)	0.17200 (17)	0.0360 (5)
C2	0.7372 (2)	0.4852 (2)	0.2855 (2)	0.0334 (6)
C3	0.8208 (2)	0.5471 (2)	0.3853 (2)	0.0410 (6)
H3	0.8945	0.5885	0.3825	0.049*
C4	0.7916 (3)	0.5456 (2)	0.4876 (2)	0.0468 (7)
H4	0.8463	0.5873	0.5568	0.056*
C5	0.6848 (3)	0.4849 (2)	0.4923 (2)	0.0501 (7)
H5	0.6678	0.4861	0.5647	0.060*
C6	0.6034 (2)	0.4233 (2)	0.3955 (2)	0.0424 (7)
H6	0.5304	0.3817	0.3998	0.051*
C7	0.6295 (2)	0.4231 (2)	0.2914 (2)	0.0356 (6)
C8	0.5662 (2)	0.3720 (2)	0.1737 (2)	0.0365 (6)
N9	0.45662 (19)	0.3047 (2)	0.1336 (2)	0.0488 (6)
C10	0.4201 (2)	0.2725 (3)	0.0174 (3)	0.0514 (8)
H10	0.3436	0.2259	−0.0154	0.062*
C11	0.4853 (3)	0.3021 (2)	−0.0574 (2)	0.0522 (8)
H11	0.4528	0.2755	−0.1384	0.063*
C12	0.5981 (2)	0.3703 (2)	−0.0164 (2)	0.0452 (7)
H12	0.6447	0.3914	−0.0664	0.054*

C13	0.6378 (2)	0.4055 (2)	0.1036 (2)	0.0370 (6)
C14	0.8358 (2)	0.5335 (2)	0.1306 (2)	0.0409 (6)
H14A	0.8610	0.6162	0.1719	0.049*
H14B	0.7963	0.5513	0.0442	0.049*
C15	0.9524 (2)	0.4548 (2)	0.1502 (2)	0.0360 (6)
C16	1.0688 (2)	0.5125 (3)	0.1743 (2)	0.0475 (7)
H16	1.0757	0.6027	0.1820	0.057*
C17	1.1752 (2)	0.4406 (3)	0.1872 (2)	0.0548 (8)
H17	1.2544	0.4817	0.2035	0.066*
C18	1.1675 (3)	0.3100 (3)	0.1769 (2)	0.0583 (8)
H18	1.2407	0.2606	0.1857	0.070*
C19	1.0529 (3)	0.2523 (3)	0.1537 (3)	0.0606 (9)
H19	1.0468	0.1621	0.1471	0.073*
C20	0.9462 (2)	0.3233 (3)	0.1399 (2)	0.0487 (7)
H20	0.8673	0.2815	0.1232	0.058*

Atomic displacement parameters (Å^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
N1	0.0301 (11)	0.0362 (12)	0.0403 (13)	−0.0008 (9)	0.0106 (9)	0.0009 (9)
C2	0.0340 (13)	0.0264 (13)	0.0380 (15)	0.0064 (11)	0.0102 (11)	0.0032 (11)
C3	0.0415 (14)	0.0330 (14)	0.0454 (17)	−0.0039 (12)	0.0113 (12)	0.0013 (12)
C4	0.0553 (17)	0.0366 (16)	0.0437 (17)	−0.0044 (14)	0.0113 (13)	−0.0011 (13)
C5	0.0628 (19)	0.0421 (16)	0.0487 (17)	0.0031 (15)	0.0234 (15)	0.0025 (14)
C6	0.0381 (15)	0.0372 (15)	0.0577 (19)	0.0002 (12)	0.0239 (14)	0.0030 (13)
C7	0.0313 (13)	0.0277 (13)	0.0459 (16)	0.0049 (11)	0.0112 (12)	0.0052 (11)
C8	0.0263 (12)	0.0303 (14)	0.0499 (16)	0.0017 (11)	0.0097 (11)	0.0029 (12)
N9	0.0354 (12)	0.0359 (13)	0.0652 (16)	0.0046 (11)	0.0050 (11)	0.0021 (12)
C10	0.0363 (15)	0.0358 (16)	0.066 (2)	0.0038 (12)	−0.0024 (15)	−0.0015 (14)
C11	0.0535 (18)	0.0396 (17)	0.0447 (17)	0.0077 (15)	−0.0066 (14)	−0.0034 (13)
C12	0.0494 (16)	0.0372 (15)	0.0442 (17)	0.0081 (13)	0.0102 (13)	0.0032 (12)
C13	0.0295 (13)	0.0314 (14)	0.0440 (16)	0.0058 (11)	0.0050 (12)	−0.0006 (12)
C14	0.0408 (14)	0.0394 (15)	0.0438 (16)	−0.0017 (12)	0.0161 (12)	0.0034 (12)
C15	0.0353 (14)	0.0397 (15)	0.0324 (14)	−0.0023 (12)	0.0112 (11)	−0.0013 (11)
C16	0.0444 (16)	0.0535 (17)	0.0422 (16)	−0.0130 (14)	0.0123 (13)	−0.0050 (13)
C17	0.0344 (15)	0.081 (2)	0.0485 (18)	−0.0076 (16)	0.0140 (13)	−0.0022 (16)
C18	0.0416 (17)	0.075 (2)	0.061 (2)	0.0133 (17)	0.0214 (14)	0.0001 (17)
C19	0.0534 (19)	0.0505 (19)	0.084 (2)	0.0045 (15)	0.0321 (17)	−0.0043 (16)
C20	0.0379 (15)	0.0436 (16)	0.0663 (19)	−0.0008 (13)	0.0202 (13)	−0.0050 (14)

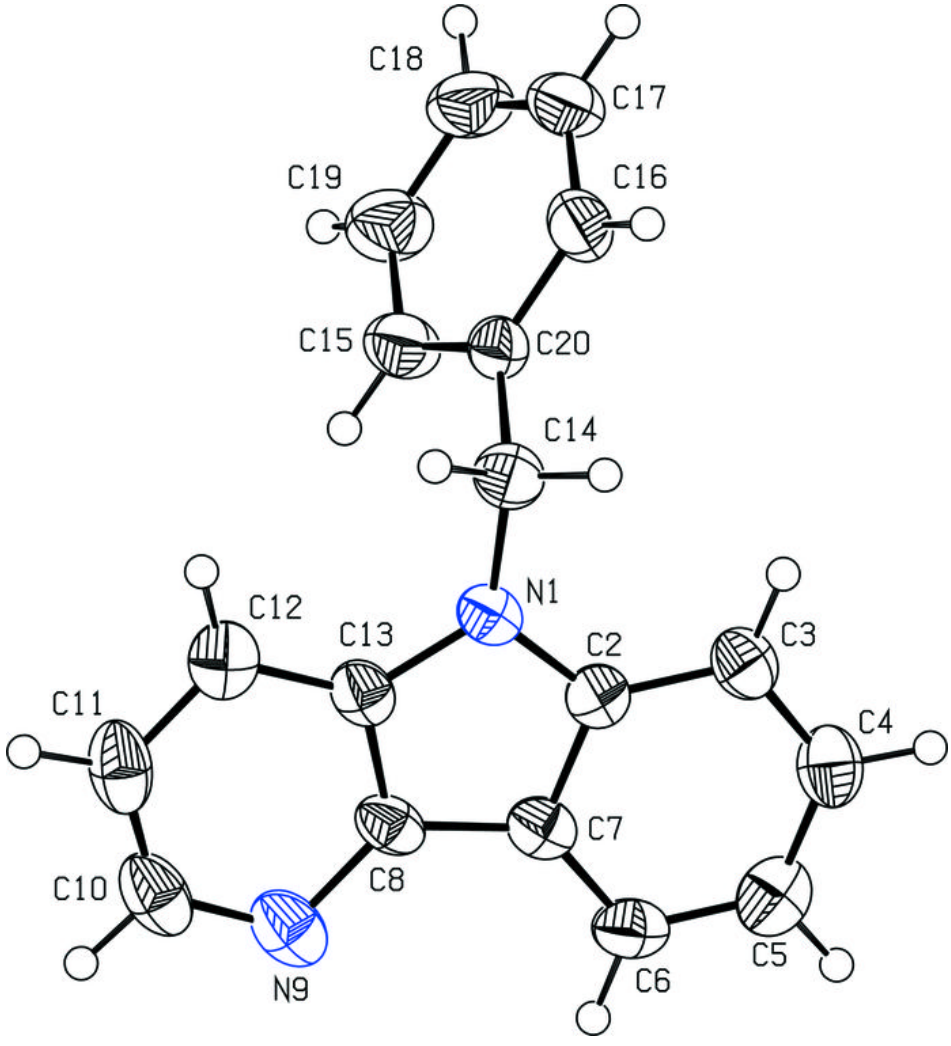
Geometric parameters (Å, °)

N1—C2	1.382 (3)	C11—C12	1.393 (4)
N1—C13	1.383 (3)	C11—H11	0.9500
N1—C14	1.453 (3)	C12—C13	1.395 (3)
C2—C3	1.397 (3)	C12—H12	0.9500
C2—C7	1.403 (3)	C14—C15	1.502 (3)
C3—C4	1.373 (3)	C14—H14A	0.9900
C3—H3	0.9500	C14—H14B	0.9900

C4—C5	1.382 (4)	C15—C16	1.383 (3)
C4—H4	0.9500	C15—C20	1.384 (3)
C5—C6	1.365 (3)	C16—C17	1.380 (4)
C5—H5	0.9500	C16—H16	0.9500
C6—C7	1.376 (3)	C17—C18	1.374 (4)
C6—H6	0.9500	C17—H17	0.9500
C7—C8	1.442 (3)	C18—C19	1.366 (4)
C8—N9	1.358 (3)	C18—H18	0.9500
C8—C13	1.398 (3)	C19—C20	1.376 (4)
N9—C10	1.348 (3)	C19—H19	0.9500
C10—C11	1.378 (4)	C20—H20	0.9500
C10—H10	0.9500		
C2—N1—C13	107.80 (19)	C11—C12—C13	115.1 (3)
C2—N1—C14	125.75 (19)	C11—C12—H12	122.4
C13—N1—C14	126.3 (2)	C13—C12—H12	122.4
N1—C2—C3	129.1 (2)	N1—C13—C12	130.5 (2)
N1—C2—C7	110.2 (2)	N1—C13—C8	109.2 (2)
C3—C2—C7	120.8 (2)	C12—C13—C8	120.3 (2)
C4—C3—C2	117.1 (2)	N1—C14—C15	114.7 (2)
C4—C3—H3	121.4	N1—C14—H14A	108.6
C2—C3—H3	121.4	C15—C14—H14A	108.6
C3—C4—C5	121.7 (2)	N1—C14—H14B	108.6
C3—C4—H4	119.1	C15—C14—H14B	108.6
C5—C4—H4	119.1	H14A—C14—H14B	107.6
C6—C5—C4	121.5 (3)	C16—C15—C20	118.0 (2)
C6—C5—H5	119.2	C16—C15—C14	120.7 (2)
C4—C5—H5	119.2	C20—C15—C14	121.2 (2)
C5—C6—C7	118.4 (2)	C17—C16—C15	120.7 (3)
C5—C6—H6	120.8	C17—C16—H16	119.6
C7—C6—H6	120.8	C15—C16—H16	119.6
C6—C7—C2	120.5 (2)	C18—C17—C16	120.6 (3)
C6—C7—C8	133.9 (2)	C18—C17—H17	119.7
C2—C7—C8	105.5 (2)	C16—C17—H17	119.7
N9—C8—C13	124.4 (2)	C19—C18—C17	119.0 (3)
N9—C8—C7	128.3 (2)	C19—C18—H18	120.5
C13—C8—C7	107.4 (2)	C17—C18—H18	120.5
C10—N9—C8	114.1 (2)	C18—C19—C20	120.9 (3)
N9—C10—C11	124.9 (3)	C18—C19—H19	119.6
N9—C10—H10	117.5	C20—C19—H19	119.6
C11—C10—H10	117.5	C19—C20—C15	120.8 (3)
C10—C11—C12	121.1 (3)	C19—C20—H20	119.6
C10—C11—H11	119.4	C15—C20—H20	119.6
C12—C11—H11	119.4		
C13—N1—C2—C3	−179.4 (2)	C10—C11—C12—C13	−0.3 (4)
C14—N1—C2—C3	−3.2 (4)	C2—N1—C13—C12	178.9 (2)
C13—N1—C2—C7	−0.1 (2)	C14—N1—C13—C12	2.8 (4)
C14—N1—C2—C7	176.1 (2)	C2—N1—C13—C8	0.1 (2)
N1—C2—C3—C4	178.4 (2)	C14—N1—C13—C8	−176.1 (2)

C7—C2—C3—C4	−0.8 (3)	C11—C12—C13—N1	−178.6 (2)
C2—C3—C4—C5	0.4 (4)	C11—C12—C13—C8	0.2 (3)
C3—C4—C5—C6	0.2 (4)	N9—C8—C13—N1	179.3 (2)
C4—C5—C6—C7	−0.3 (4)	C7—C8—C13—N1	0.0 (3)
C5—C6—C7—C2	−0.2 (4)	N9—C8—C13—C12	0.3 (4)
C5—C6—C7—C8	−178.6 (2)	C7—C8—C13—C12	−179.0 (2)
N1—C2—C7—C6	−178.6 (2)	C2—N1—C14—C15	88.4 (3)
C3—C2—C7—C6	0.7 (4)	C13—N1—C14—C15	−96.1 (3)
N1—C2—C7—C8	0.1 (2)	N1—C14—C15—C16	−146.8 (2)
C3—C2—C7—C8	179.5 (2)	N1—C14—C15—C20	35.9 (3)
C6—C7—C8—N9	−0.8 (4)	C20—C15—C16—C17	0.2 (4)
C2—C7—C8—N9	−179.3 (2)	C14—C15—C16—C17	−177.2 (2)
C6—C7—C8—C13	178.4 (3)	C15—C16—C17—C18	−0.3 (4)
C2—C7—C8—C13	−0.1 (2)	C16—C17—C18—C19	−0.1 (4)
C13—C8—N9—C10	−0.6 (3)	C17—C18—C19—C20	0.5 (4)
C7—C8—N9—C10	178.5 (2)	C18—C19—C20—C15	−0.5 (4)
C8—N9—C10—C11	0.5 (4)	C16—C15—C20—C19	0.2 (4)
N9—C10—C11—C12	−0.1 (4)	C14—C15—C20—C19	177.6 (2)

Fig. 1


C18
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